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In addition, mass and molecular weight will give us moles. It appears that the ideal

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gas law is called for. However, there is a problem. We are being asked to change the conditions to a new amount of moles and pressure. So, it seems like the ideal gas law needs to be used twice. 2) Let's set up two ideal gas law equations: $P_1 V_1 = n_1 RT_1$

ChemTeam: Ideal Gas Law: Problems

#1 - 10
Page 5/25

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2) At what temperature would 2.10 moles of N_2 gas have a pressure of 1.25 atm and in a 25.0 L tank? 3) When filling a weather balloon with gas you have to consider that the gas will expand greatly as it rises and the pressure decreases.

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Let's say you put about 10.0 moles of He gas into a balloon that can inflate to hold 5000.0L. Currently,

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Holt ChemFile: Problem-
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Additional Problem Solving

Solving Workbook 191

Stoichiometry of Gases

Name Class Date

Problem Solving

continued Rearrange

the ideal-gas-law

equation to solve for
the unknown quantity,

V . $PV = nRT$ $V = \frac{nRT}{P}$

COMPUTE EVALUATE

Are the units correct?

Yes; units canceled to
give liters of SO_2 . Is

the number of
significant figures
correct?

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Workbook 191
Stoichiometry ...**

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**Home - Crestwood
Local School District**

Ideal Gas Law

Worksheet $PV = nRT$

Use the ideal gas law,
“ $PV = nRT$ ”, and the
universal gas constant
 $R = 0.0821 \text{ L}\cdot\text{atm}$ to
solve the following
problems: $\text{K}\cdot\text{mol}$ If

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pressure is needed in kPa then convert by multiplying by $101.3 \text{ kPa} / 1 \text{ atm}$ to get $R = 8.31 \text{ kPa} \cdot \text{L} / (\text{K} \cdot \text{mole})$

1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 ...

Ideal Gas Law

Worksheet PV = nRT

The Ideal Gas Law, continued
The Ideal Gas Law Relates All Four Gas Variables, continued

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Problems for
problems that use units of kilopascals and liters when using the ideal gas law, the value you will use for R is as follows: Section 3 Molecular Composition of Chapter 12 Gases • If the pressure is expressed in atmospheres, then the value of R is:

The Ideal Gas Law - crestwoodschools.org

When the wall that

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Problem 10.11
separates the two chambers within the tank is removed, the air expands to fill the right side of the tank. Calculate the final temperature and pressure in the tank.: Assume air behaves as an ideal gas and the process is adiabatic because the tank is well-insulated.: Read : The most important thing to recognize in this problem is that removing the partition

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is equivalent to ...
Problems Holt

**Example 5E - 3:
Expansion of an
Ideal Gas to Fill an**

...

Ideal Gas Law

Worksheet $PV = nRT$

Use the ideal gas law,

" $PV = nRT$ ", and the

universal gas constant

$R = 0.0821 \text{ L}\cdot\text{atm}$ to

solve the following

problems: $\text{K}\cdot\text{mol}$ If

pressure is needed in

kPa then convert by

multiplying by

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101.3kPa / 1atm to get
R = 8.31 L*kPa /
(K*mole) 1) If I have 4
moles of a gas at a
pressure of 5.6 atm
and a volume of 12
liters ...

Ideal Gas Law

Worksheet PV = nRT

- Quia

the gas is measured at
STP, you will need only
Avogadro's law to
relate the volume and
amount of a gas. One
mole of any gas at STP

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occupies 22.4 L. If the gas is not at STP, you will need to use the ideal gas law to determine the number of moles. Once volume has been converted to amount in moles you can use the mole ratios of products and ...

CHEMFILE MINI- GUIDE TO PROBLEM SOLVING CHAPTER 13 ...

The ideal gas law is an equation of state the

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Problem Set

describes the behavior of an ideal gas and also a real gas under conditions of ordinary temperature and low pressure. This is one of the most useful gas laws to know because it can be used to find pressure, volume, number of moles, or temperature of a gas.

Ideal Gas Law Example Problem - ThoughtCo

The ideal gas law

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relates the pressure, volume, quantity, and temperature of an ideal gas. At ordinary temperatures, you can use the ideal gas law to approximate the behavior of real gases. Here are examples of how to use the ideal gas law. You may wish to refer to the general properties of gases to review concepts and formulae related to ideal ...

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Additional **Ideal Gas Law: Worked Chemistry Problems - ThoughtCo**

(Alternatively, consider a gas of N non-identical particles.) Find and sketch the heat capacity as a function of temperature. 6.28 Pressure and energy density (This problem was inspired by Reif problem 9.5.) Any non-relativistic monatomic ideal gas, whether classical or quantal,

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satisfies \[

$$p = \frac{2}{3}$$

$$\frac{E}{V}.\]$$

6.9: Additional Problems - Physics LibreTexts

The Named Gas Laws
"derived" using the
Ideal Gas Law.

Additional Gas Laws

Combined Gas Law

(usually PVT vary, n
constant) Ideal Gas

Law ($PV = nRT$)

Dalton's Law of Partial

Pressures; Graham's

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Additional
Law of Effusion.

Additional Gas-related

Tutorials & Problems

Molar Volume; Gas

Density; Vapor

Pressure; Gas Velocity;

The Clasius-Clapeyron

Equation

ChemTeam: KMT & Gas Laws

' The volume of a given mass of an ideal gas is directly proportional to the temperature in kelvin and number of moles (or molecules)

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and inversely
proportional to the
pressure.

(PDF) Worked Examples on Gas Laws and Kinetic Theory

Holt ChemFile: Problem-
Solving Workbook 51

Mole Concept Name

Class Date Problem

Solving continued

Sample Problem 2 A

student needs 0.366

mol of zinc for a

reaction. What mass of

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zinc in grams should
the student obtain?

Solution ANALYZE

What is given in the
problem? amount of
zinc needed in moles

What are you asked to
find? mass of zinc in
grams ...

Copyright code: d41d8
cd98f00b204e9800998
ecf8427e.

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