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set (49) Valence electrons. The electrons that have the highest energy in an atom. They are involved in chemical bonding. The number of valence electrons in each atom helps determine the chemical ...

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Multiple Choice Write the letter of the

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correct answer on the line at the left.

___ 1. Which is a property shared by most molecular compounds? a. high boiling point b. high melting point c. low melting point d. nonpolar bonds ___ 2.

When an atom loses an electron, it becomes a a. positive ion. b. negative ion.

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Atoms and Bonding - Bridgeway

Atoms and Bonding Chapter Tests Atoms and Bonding (continued) 30. A metal crystal consists of positively charged metal ions embedded in a “sea” of loosely held valence electrons that can move around easily. Heat travels through materials as the increased motion of the particles in the hotter

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parts of the material is passed along to the

Answers Atoms and Bonding Group 2 element.

Atoms and Bonding Chapter Test A
Using Science Skills: Interpreting
Diagrams Use the electron dot diagrams to answer the following. 21. Applying

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Concepts Predict the formula for the compound, if any, that would form from the elements listed below. If a compound is unlikely, explain the reason why.

a. magnesium and chlorine

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bonniboo. Terms in this set (67) a
charged group of covalently bonded
atoms that tends to act like a single
atom is a. polyatomic atom. a charged
particle is referred to as. an ion.

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Chapter 5 Atoms and Bonding Chapter Preview Questions 4. Compounds are formed by a. combining two or more different elements. b. bombarding atoms with high-speed particles. c. combining two or more different nuclei. d. dissolving a solid in a liquid.

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science 4.1 4.2 4.3 17 Terms.
sturiskyryan.

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What kind of bonds do these compounds have? a. ionic bonds b. covalent bonds c. metallic bonds ____ 24. Applying Concepts Which compound has a polar

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bond? a. HF b. F₂ c. O₂ ____ 25.

Applying Concepts Which compound has a triple bond? a. O₂ b. F₂ c. N₂ H (+)
(-) F O O N N F F Atoms and Bonding

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